HW04 - Electromagnetic Radiation

⚠ This is a preview of the draft version of the quiz

Started: Aug 8 at 4:50pm

Quiz Instructions

Homework 04 - Electromagnetic Radiation

Question 1 **1 pts**

What is the frequency of light with a wavelength of 4.0×10^{-7} m?

- $1.3 \times 10^{-15}\ s^{-1}$
- $3.0 \times 10^{-14}\ s^{-1}$
- $3.0 \times 10^{14}\ s^{-1}$
- $7.5 \times 10^{14}\ s^{-1}$

Question 2 **1 pts**

What is the correct order of increasing frequency?

- radio waves, visible light, ultraviolet radiation, infrared radiation
- radio waves, infrared radiation, ultraviolet radiation, visible light
- radio waves, infrared radiation, visible light, ultraviolet radiation
- infrared radiation, radio waves, visible light, ultraviolet radiation
- ultraviolet radiation, visible light, infrared radiation, radio waves

Question 3 **1 pts**

Light with a frequency of 7.30×10^{14} Hz lies in the violet region of the visible spectrum. What is the wavelength of this frequency of light?

- 411 nm
- 4.11×10^{-7} nm
- 4.11×10^{21} nm
- 4.11×10^{-16} nm

Question 4 **1 pts**

When an electron beam strikes a block of copper, x-rays of frequency 1.97×10^{19} Hz are emitted. What is the wavelength of these x-rays?

- 1.52×10^{-11} nm
- 15.2 nm
- 1.52×10^{-2} pm
- 15.2 pm

Question 5 **1 pts**

Wavelength is...

- the distance between successive peaks in a wave.
- one-half of the height of a wave.
- the distance between a peak of one wave and the trough of the next.
- the number of waves passing a fixed point in one second.

Question 6 **1 pts**

Frequency is...

- the distance between successive peaks in a wave.
- the distance between a peak in one wave to the trough in the next wave.
- one half the height of the wave.
- the number of waves passing a fixed point in one second.

Question 7 **1 pts**

It takes light with a wavelength of 212 nm to break the N–H bond in ammonia. What energy is required per photon to break this bond? What is the N–H bond strength in terms of kJ per mole?

- 9.4×10^{-22} kJ/photon; 565,000 kJ/mol
- 6.6×10^{-22} kJ/photon; 398 kJ/mol
- 9.4×10^{-22} kJ/photon; 565 kJ/mol
- 6.6×10^{-22} kJ/photon; 0.398 kJ/mol

Question 8 **1 pts**

In 1 sec, a 60 W bulb emits 11 J of energy in the form of infrared radiation (heat) of a corresponding wavelength of 1850 nm. How many photons of infrared radiation does the lamp generate in 1 sec?

- 1.04×10^{29} photons
- 1.10×10^{-19} photons
- 6.63×10^{23} photons
- 1.02×10^{20} photons

Question 9 **1 pts**

A photon has a frequency of 223 MHz. What is the energy of this photon?

- 1.48×10^{-25} J
- 1.48×10^{-31} J
- 8.91×10^{-22} J
- 8.91×10^{-28} J

Question 10 **1 pts**

Carbon emits photons at 745 nm when exposed to blackbody radiation. How much energy would be obtained if 44g of carbon were irradiated? Assume each carbon atom emits one photon.

- 9.1×10^{5} J
- 2.7×10^{-19} J
- 7.1×10^{6} J
- 5.9×10^{5} J

Question 11 **1 pts**

A 200 nm photon has _____ times the energy of a 700 nm photon.

- 4.2
- 0.29
- 3.5
- 0.37

Question 12 **1 pts**

If a photon's wavelength is 663 μm, what is its energy?

- 3.00×10^{-22} J
- 4.40×10^{-46} J
- 3.00×10^{-25} J
- 4.40×10^{-43} J

Question 13 **1 pts**

Sodium vapor lamps, used for public lighting, emit yellow light of a wavelength of 570 nm. How much energy is emitted by an excited sodium atom when it generates a photon?

- 3.5×10^{-19} J
- 2.8×10^{-19} J
- 3.5×10^{-28} J
- 2.8×10^{-20} J

Question 14 **1 pts**

Consider the sodium vapor lamps described in question 13. How much energy is emitted by 45.8 mg of sodium atoms emitting light at this wavelength? Assume each sodium atom emits one photon.

- 420 J
- 2.0×10^{21} J
- 2.0×10^{-3} J
- 4.2×10^{5} J

Question 15 **1 pts**

A particular metal has a work function of 1.05 eV. A light is shined onto this metal with a corresponding wavelength of 324 nm. What is the maximum velocity of the photoelectrons produced? (Hint: 1eV = 1.6022×10^{-19} J, mass of an electron = 9.11×10^{-31} kg)

- No photoelectrons are produced.
- 1.35×10^{12} m/s
- 9.89×10^{5} m/s
- 1.16×10^{6} m/s

Question 16 **1 pts**

A particular metal has a work function of 3.05 eV. A light is shined onto this metal with a corresponding wavelength of 524 nm. What is the maximum velocity of the photoelectrons produced? (Hint: 1eV = 1.6022 x 10^{-19} J, mass of an electron = 9.11 x 10^{-31} kg)

- 8.72 x 10^{8} m/s
- 9.12 x 10^{5} m/s
- 8.32 x 10^{11} m/s
- No photoelectrons are produced.

Question 17 **1 pts**

Max Planck's theory averted the so called "UV Catastrophe" of classical mechanics. Which of the following best describes *how* Planck's theory avoided the "UV Catastrophe"?

- Radiation given off by blackbody radiators can be emitted in all types of radiation, not just UV radiation.
- Radiation emitted by blackbody radiators will reach UV energy levels only at extremely high temperatures.
- Eventually, blackbody radiators can cool to a temperature of absolute zero, resulting in its inability to release any more UV radiation.
- Radiation given off by blackbody radiators can only be emitted in quantized amounts.

Question 18 **1 pts**

The de Broglie equation was important for a number of reasons, not least of which was that it demonstrated that _____.

- all objects have a wavelength. However, in the case of macroscopic objects, these wavelengths are so small that they can be ignored.
- only quantum objects have wavelengths.
- all objects have a wavelength. However, in the case of quantum objects, these wavelengths are so small that they can be ignored.
- only macroscopic objects have wavelengths.

Question 19 **1 pts**

An atom of which element, moving at 240 m/s, would possess a de Broglie wavelength of 1.40×10^{-11} m?

- Sn
- Mn
- Cs
- At

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